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Solutions Manual Chemistry: Matter and Change • Chapter 11
211 CHAPTER 11 SOLUTIONS MANUAL Section 11.1 Assessment
page 372 5. Compare the mass of the reactants and the mass of the products in a chemical reaction, and explain how these masses are related. The coefficients in the balanced equation

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163 CHAPTER 10 SOLUTIONS MANUAL 3. Apply How many atoms of hydrogen-1 are in a 1.007-g sample? Recall that 1.007 amu is the mass of one atom of hydrogen-1. Round your answer to two significant digits. 1.007 g 1 H-1 atom $\underline{\quad}$ 1.672 10²⁴ g a. 6.023 10²³ H-1 atoms 4.

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212 Chemistry: Matter and Change • Chapter 11 Solutions Manual CHAPTER 11 SOLUTIONS MANUAL c. How many moles of H_2S are produced? $1.50 \text{ mol S} \times \frac{4 \text{ mol H}_2\text{S}}{8 \text{ mol S}} = 0.75 \text{ mol H}_2\text{S}$ 12. Challenge Sulfuric acid (H_2SO_4) is formed when sulfur dioxide (SO_2) reacts with oxygen and water. a. Write the balanced chemical equation for the reaction. $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{SO}_4(\text{aq})$

StoichiometryStoichiometry

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